



Collision Theory

• For a chemical reaction to occur, reactants must :

a)

b)

• Rate of Reaction can be increased by: a)

b)

c)

Increasing Rate

↑ Concentration

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•

↑ Surface Area(s)

•

•

↑ Temperature

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•

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+ Catalyst

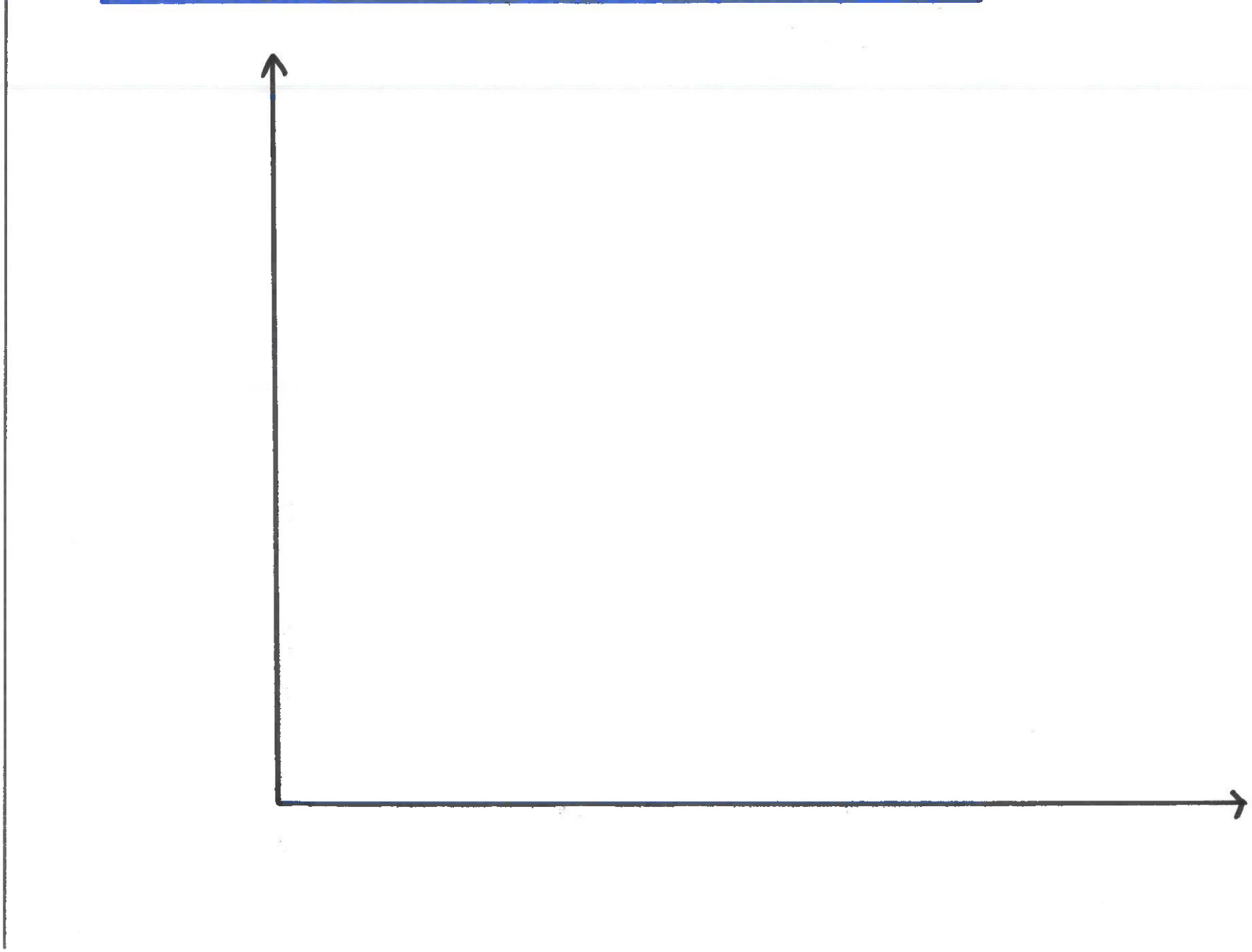
•

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Maxwell-Boltzmann Distribution Curve

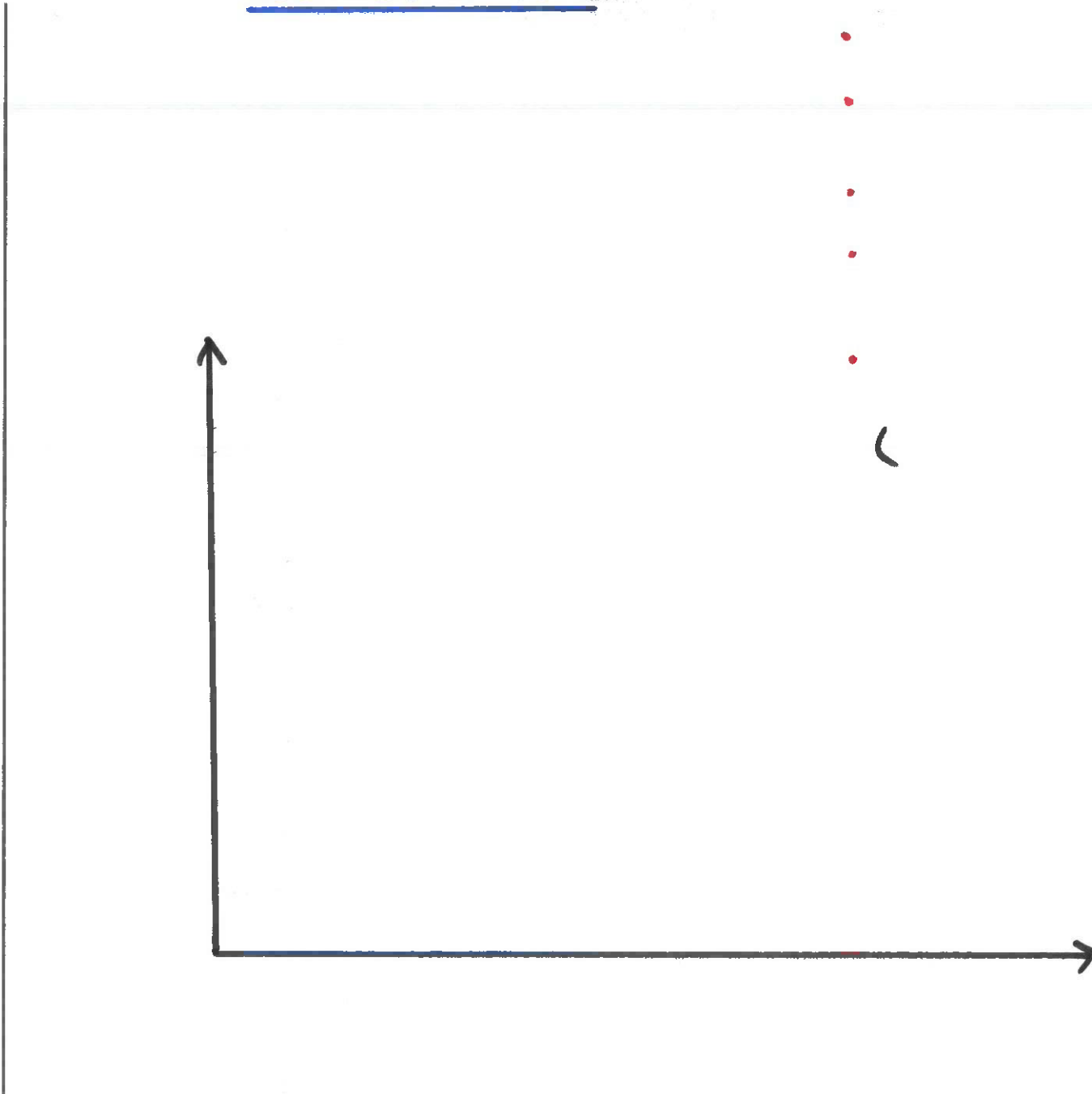
- 1
- 2
- 3
- 4
- 5





Concentration

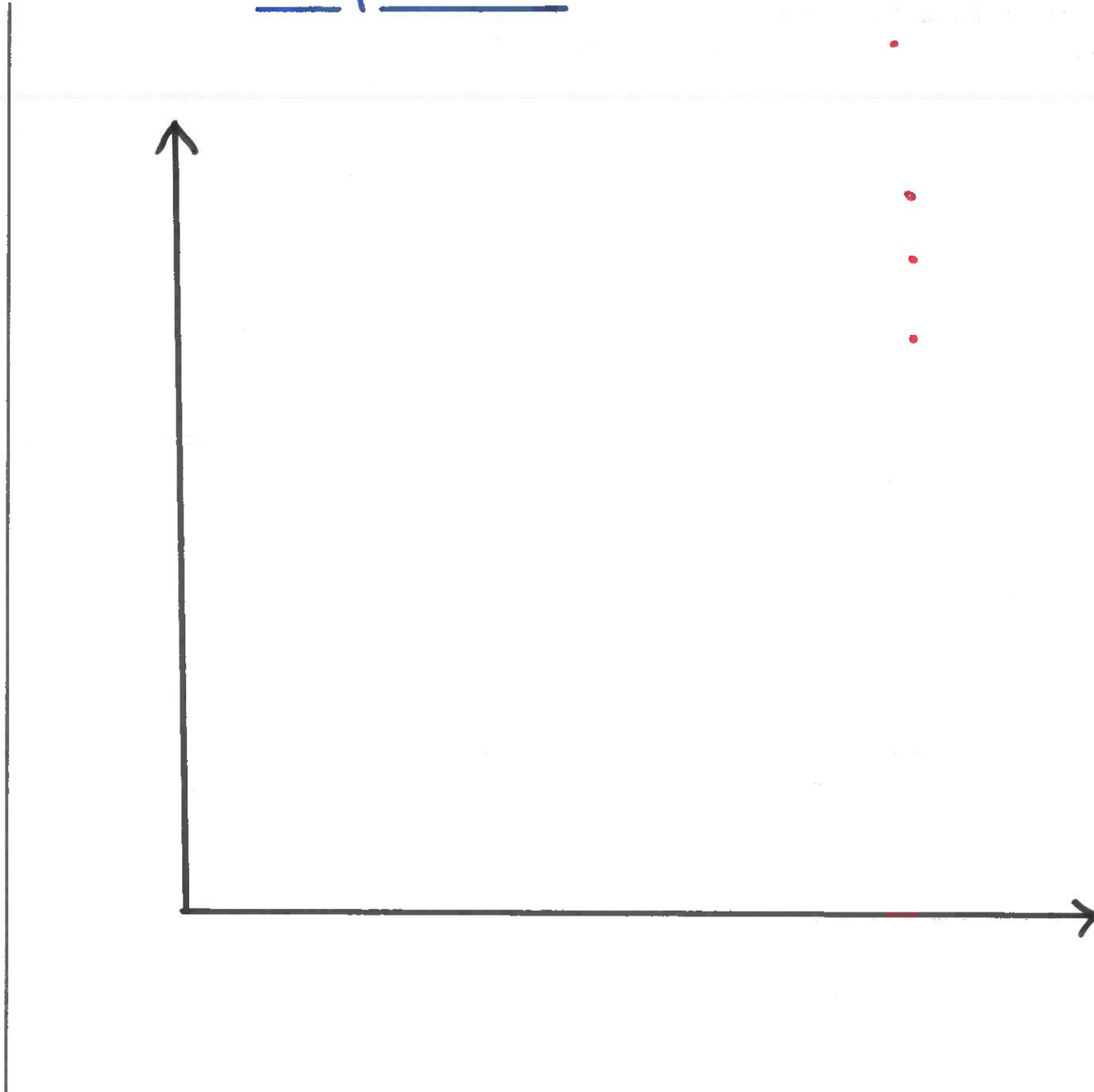
- 1
- 2
- 3
- 4
- 5





Temperature

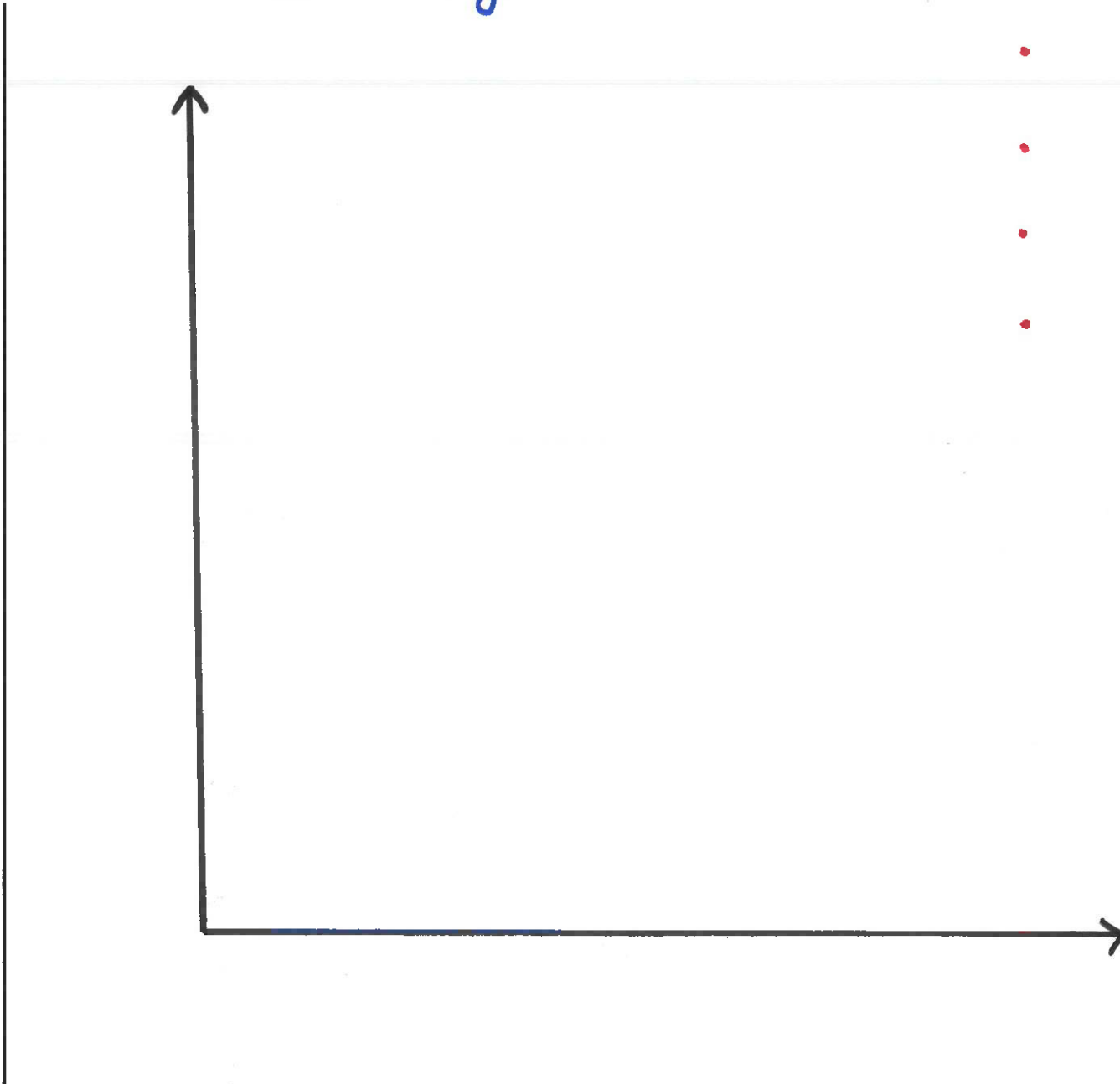
- 1
- 2
- 3
- 4
- 5





Catalyst

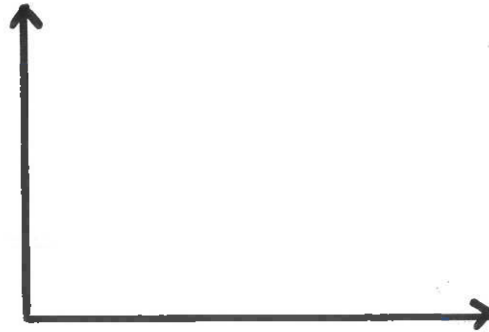
- 1
- 2
- 3
- 4
- 5





How Catalysts Work

How do Catalysts work?



Homogenous Catalyst

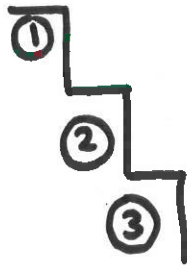
Homogeneous Catalyst

e.g.

Heterogeneous Catalyst

Heterogenous Catalyst

e.g.



Examples

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